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# Colloidal and Thermal Behaviors of Some Venezuelan Kaolin Pastes for Therapeutic Applications

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**Abstract:** This work contributes to the knowledge of colloidal and thermal properties of some important primary-originated kaolin deposits from Bolivar State, Venezuela, and their possible use as semisolid formulations in medicinal muds for topical applications. Eight selected high purity kaolin samples were characterized. Rheological and thermal properties were correlated to physico-chemical characteristics of the clay suspensions (pH, Ca<sup>2+</sup>, and Mg<sup>2+</sup> cation desorption and surface charge). Most of the studied kaolin pastes showed adequate viscosities, acceptable skin safe pH, and good thermal properties for pelotherapeutic uses. Three of the studied samples, in particular, showed very high kaolinite purities (>92% kaolinite), elevated viscosities (>1 Pa·s), and good thermal and pH performances for topical application.

**Keywords:** kaolin; mud therapy; thermal properties; zeta potential; rheology

## 1. Introduction

Clay minerals have been widely used to develop products and functionalities in a large variety of areas, including materials, pharmaceuticals, food sciences, and biotechnology. Mixtures of clays and mineral waters are used in health mud therapies, either for thermotherapeutic or cosmetic purposes, and constitute the basis of most of the muds used in healing and spa centers. The ever-increasing use of clay minerals in pelotherapeutic applications is due to the versatility and peculiarity of their physical and chemical properties and the suitability of their colloidal and thermal behaviors [1–15]. Pioneer studies [16,17] indicated that rheology is mainly controlled by the morphology and aggregation state of solid particles, as well as the nature and strength of the predominant interparticle electrostatic forces (i.e., as predicted by zeta potential measurements), which are fundamental parameters to qualify the properties of clayey dispersions, such as spreading qualities, adhesion to skin and removal capacities [16,18–22], and these are affected by pH, concentration, and relative dimensions of solids [23].

On the other hand, clay thermal behaviors (cooling kinetics and heat capacity) are directly responsible for their therapeutic effect [3,5,24–26].

Kaolinite is the main clay mineral of the kaolin group, and it has attracted great interest in pharmaceuticals and therapy fields based on their properties [27]. Technologically, the optimal water content for the formulation of consistent and adequate adhesive kaolinite pastes is a 1:1 (*w/w*) solid/liquid ratio of the prepared aqueous dispersions [28]. In addition to the influence of water content and solid fraction composition in the quality and healing activity of the thermotherapeutic muds, Awad et al. [29] proved that the thermal efficiency and performance of therapeutic kaolin pastes (heat retention capacity and prolonged release) are controlled by the granulometric properties of kaolin powder (i.e., including the particle size, grading, surface area, and particle textural packing), and the degree of structural order-disorder and crystallite size of the kaolinite. As the appropriate uniformity grading and particle fineness of the kaolin powder were found to be controlled by the mineral's structural characteristics, the later ones correlate directly with its composition and chemical variability [30].

Concerning economically viable Venezuelan kaolin resources, Hernández et al. [31] characterized their compositional, structural, and textural characteristics in order to evaluate their suitability as health care components. On the basis of these results, the aim of the present work is to determine the colloidal and thermal properties of representative high-grade Venezuelan kaolin samples to evaluate their possible use as semisolid formulations in medicinal peloids.

## 2. Materials and Methods

### 2.1. Materials and Paste Formulation

Eight representative high-grade kaolin samples (collected from different primary kaolin deposits on the Guayana shield, in Bolivar State, Venezuela), namely VE1, VE2, VE3, VE4, VE5, VE6, VE9, and VE12, were selected for the present study. Their geological context and mineralogical compositions were described in Hernández et al. [31]. Each sample was dried, pulverized, and sieved to obtain particle aggregates below 125  $\mu\text{m}$ . Kaolinite pastes 1:1 (*w/w*) were prepared with purified water (at 25 °C ambient temperature) using a vortex mixer and then kept moist in a refrigerator prior to the targeted characterizations.

### 2.2. pH Measurement

The pH of the aqueous suspensions of kaolin was measured under mechanical agitation using a Crison pH-meter Basic 20+ equipped with a semisolid sensor (5053 T) calibrated by standard solutions at pH 4, 7, and 9. The pH tolerance ranges between 2 and 11. Three replicates were carried out for each measurement and results averaged ( $\pm 0.1$  standard deviation).

### 2.3. Apparent Viscosity Measurement

Rheological analysis was carried out by a computerized controlled rate viscometer (RotoVisco 1; HAAKE RheoWin software, Thermo Scientific HAAKE, Waltham, MA, USA) as a triple measure of apparent viscosity ( $\eta$  in Pa·s) at a controlled temperature (25 °C) after a rest time of 90 s in the shear rate range 10–800  $\text{s}^{-1}$ . Before treatment, samples were sonicated for 3 min using Power Sonic 405 ultrasonic equipment (Hwashin, Tech Co. Ltd., Gangnam-gu, Seoul, Korea). Four replicates were performed for each sample.

### 2.4. Zeta Potential Measurement

Zeta potential measurements were performed using an electrophoretic light scattering (ELS) Zetasizer Nano-Z<sup>TM</sup> (Malvern Instruments, Malvern, UK). The measurements were taken through a downward ramp from pH 3 to pH 9. The resultant suspensions (10 mg of powder sonicated in 100 mL of deionized water) were introduced into folded 1 mL capacity capillary zeta potential cells (DTS1061, Malvern Instruments). Measurements were taken at a temperature of  $25 \pm 0.1$  °C, and three replicates

were obtained for each sample. The zeta potentials and iso-electric points at each corresponding pH condition were performed using Smoluchowski's equation [32].

### 2.5. Desorbed Cation Measurements

Experimental cation desorption was performed following the same method described in Sánchez-Espejo [15]. The aqueous supernatant phases were separated from the solid mud by centrifugation and filtration through a membrane filter with a 0.05  $\mu\text{m}$  pore size.  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  cations in the filtrates were then analyzed by means of ICP-OES technique.

### 2.6. Thermal Measurements

Specific heat was determined by means of a differential scanning calorimetry (DSC) technique using a Shimadzu<sup>®</sup> DSC-50 Calorimeter (Shimadzu Corporation, Nakagyo Ward, Kyoto, Japan). Samples were introduced in a closed 50 mL cylindrical Teflon container and heated at constant temperature (70 °C) and immersed in a thermostatic bath (Edelstahl Rostfrei<sup>®</sup>) at 35 °C. The temperature was measured every 30 s by a Dual Thermometer LT Lutron TM-906A until the temperatures of the kaolin paste sample and the thermostatic bath were equal. Cooling curves were optimized following the procedure described by Sánchez-Espejo et al. [15], where cooling kinetics were calculated following Cara et al. [3]. Experimental cooling data were fitted using Newton's Law, describing the thermal exchange between two bodies in contact at different temperatures:

$$(T - T_{\min}) = (T_{\max} - T_{\min}) e^{-Kt} \quad (1)$$

where,  $T_{\min}$  was the room temperature (25 °C),  $T_{\max}$  was the initial temperature (50 °C),  $t$  was the time in minutes, and  $K$  was the constant that depends on the material and apparatus, given by:

$$K = \frac{P}{C} = \frac{P}{mC_p} \quad (2)$$

where,  $P$  is the instrumental constant of the apparatus,  $C$  the heat capacity of the heated material,  $m$  the heated mass, and  $C_p$  the specific heat. Experimental thermal parameters of the studied samples were then obtained using the aforementioned equations.

## 3. Results and Discussion

### 3.1. Mineralogy

Table 1 summarizes the mineral composition and Hinckley Index (HI) of the studied samples, determined by X-ray diffraction and chemical analysis [31]. Samples essentially consisted of kaolinite, with contents ranging from 73% to 99%, with an average of 88.3%. Other detected minerals included quartz, gibbsite, hematite, anatase, rutile, and mica as impurities. According to the HI, kaolinite in all samples had a basically ordered structure (HI > 0.7). Contents of trace elements in the global kaolin deposits were attributed to their substitutions in the crystal structures of the associated titanium oxides and iron oxy-hydroxide minerals. Hence, the Pb and As (ppm) contents can be easily reduced to safe limits by removal of these impurities following reported kaolin purification methods [27].

**Table 1.** Kaolinite contents, degree of structural order-disorder (HI), mineral impurities, iron oxide contents, and desorbed cations in the formulated kaolin samples.

Samples	Kaolinite (%)	HI	Associated Mineral Impurities	Fe <sub>2</sub> O <sub>3</sub> (%)	Desorbed Cations (mg/L)	
					Ca <sup>2+</sup>	Mg <sup>2+</sup>
VE1	75	1.2	Gibbsite (20%), Fe–Ti oxide minerals (4%), quartz (1%)	1.79	29.8	8.59
VE2	92	0.7	Gibbsite (4%), Fe–Ti oxide minerals (3%), quartz (1%)	0.70	70.57	75.08
VE3	94	1.3	Gibbsite (5%), anatase (1%)	0.34	26.33	46.92
VE4	99	1.2	Gibbsite (1%).	0.37	0.77	-
VE5	88	1.4	Mica (5%), quartz (6%), anatase (1%)	0.37	-	-
VE6	92	1.4	Gibbsite (5%), Fe–Ti oxide minerals (2%), quartz (1%)	2.06	4.61	-
VE9	94	0.9	Quartz (4%), Fe–Ti oxide minerals (2%)	0.81	-	-
VE12	73	0.9	Gibbsite (13%), quartz (3%), zircon (2%), Fe–Ti oxide minerals (9%).	7.76	155.80	47.08

### 3.2. Rheology

Table 2 summarizes the apparent viscosity of the studied water-based kaolin pastes, calculated at a shear value of  $250 \text{ s}^{-1}$ , at  $25 \text{ }^\circ\text{C}$ , which is representative of the stress produced by skin spreading during applications [33]. The aqueous dispersions of kaolin showed typical non-Newtonian viscoplastic flow curves with variable behavior. VE9 showed the highest average apparent viscosity ( $3.96 \text{ Pa}\cdot\text{s}$ ), while VE1, VE3, VE4, and VE6 also showed high apparent viscosity ( $>1 \text{ Pa}\cdot\text{s}$ ). Possibly as a result of the significant iron oxide contents ( $7.76\%$ ), the dispersions formulated with VE12 showed negligible apparent viscosity ( $0.03 \text{ Pa}\cdot\text{s}$ ), which would prevent the permanence of the product during application. In general, the viscosities measured in these Venezuelan kaolins are high in comparison to other kaolins [34–37]. Studies on other types of clay used for pelotherapy (bentonites) showed viscosities higher than  $1 \text{ Pa}\cdot\text{s}$  [38,39]. Thus, the higher values of viscosity in the kaolin suspensions would facilitate management and permanence on the application area of the skin, whereas lower values tend to have inadequate consistency [2,10,40]. Therefore, dispersions formulated with VE6 and VE9 had the best potential to form clay pastes with rheological behavior adequate for treatments.

**Table 2.** Viscosity (at  $250 \text{ s}^{-1}$ ,  $25 \text{ }^\circ\text{C}$ ), pH values, and the median size ( $D_{50}$ ) of the studied kaolin dispersions.

Samples	Flow Character	Viscosity (Pa·s) $\pm$ S.D. ( $n = 4$ )	Suspension pH $\pm$ S.D. ( $n = 3$ )	$D_{50}$ ( $\mu\text{m}$ )
VE1	Pseudoplastic	$1.82 \pm 0.10$	$4.30 \pm 0.01$	1.05
VE2	Dilatant	$0.70 \pm 0.05$	$4.50 \pm 0.02$	1.12
VE3	Dilatant	$1.13 \pm 0.30$	$4.40 \pm 0.07$	1.24
VE4	Dilatant	$1.44 \pm 0.20$	$4.52 \pm 0.09$	1.55
VE5	Dilatant	$0.32 \pm 0.10$	$4.50 \pm 0.04$	2.29
VE6	Dilatant	$2.98 \pm 0.20$	$4.70 \pm 0.05$	5.1
VE9	Pseudoplastic	$3.96 \pm 0.20$	$5.90 \pm 0.12$	1.81
VE12	Dilatant	$0.03 \pm 0.10$	$3.86 \pm 0.89$	1.83

The rheological behavior of kaolin suspensions is governed by size, morphology, aggregation, and surface charge in the clay mineral particles [16,17]. Favorable rheological behaviors are expected with a combination of coarse, medium, and fine grains because interferences are minimized due to the reduction of the free space between the rotation spheres of bigger particles [41]. The studied kaolins exhibited highly variable granulometric  $D_{50}$  and  $D_{90}$  values (from approximately  $1$  to  $5 \mu\text{m}$  and  $4$  to

52  $\mu\text{m}$ , respectively, [31]). Moreover, samples VE1, VE3, VE4, VE6, and VE9, with  $> 1$  Paviscosities, present at least two groups with different kaolinite particle sizes [31]. On the other hand, the presence of specific impurities modified the resultant viscosity. VE12, with the highest amount of impurities (Table 1), exhibited the lowest viscosity values (Table 2). The presence of mica reduced the viscosity in VE5, whereas the presence of gibbsite [42] increased the viscosities and produced shear thickening flows in those samples with noticeable amounts of this mineral.

### 3.3. pH

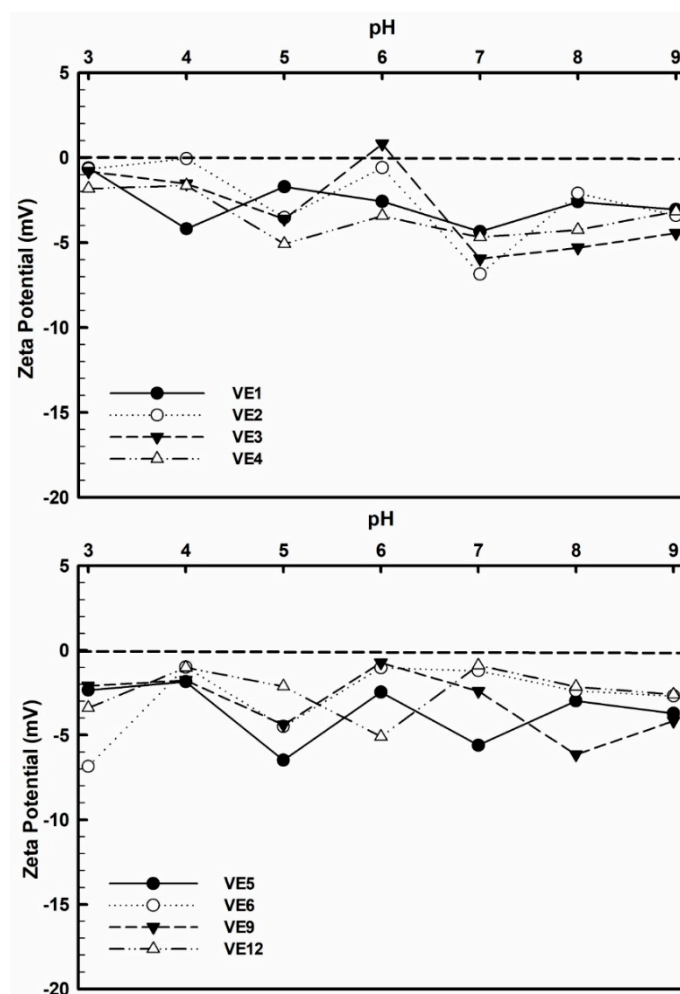
Table 2 also summarizes the pH values of the studied kaolins dispersed in purified water. The samples were acid in character, with pH ranges from 3.86 to 5.97. The pure Fe–Ti oxide minerals and gibbsite were stable at  $\text{pH} > 5$ , while they could be dissociated under  $\text{pH} < 5$  conditions. However, the highest  $\text{Ca}^{2+}$  content (155.80 mg/L) found in sample VE12, with its variable mineralogical composition, probably had the most influence on the ionic strength of the kaolin primary colloidal solutions and hence on the measured aqueous dispersions, which could maintain stabilities of the associated gibbsite and Fe–Ti oxide mineral phases till  $\text{pH} \approx 4$ . The pH values of the dispersions prepared with samples VE2, VE3, VE5, VE6, and VE9 could permit topical application in terms of the pH tolerability range of the human skin (4.5–6). Specifically, the pH of VE6 was closer to the most healthy human skin pH (approximately 4.7), which is favorable to skin microflora [43] and prevents the growth of pathogenic bacteria [44]. It has been reported that  $\text{Al}(\text{OH})_3$  (gibbsite) is insoluble at pH 5 to 8, but the dissolution of Al increases under more acid conditions [45]. However, clinical trials showed that the aluminum compounds, as used in cosmetic products, are dermatologically safe, as they have no toxic effects and do not irritate the skin [46,47]. Moreover, the acidic character of the studied kaolin-based pastes can be optimized with pharmaceutical pH buffer mediators.

### 3.4. Zeta Potential

Figure 1 shows the zeta potential (ZP) of the studied kaolins at their corresponding pH values. The measured ZP profiles were mainly steady and oscillated slightly around  $-5$  mV within the pH 3–9 range. ZP values were negative due to the permanent negative charges on the kaolinite particle surfaces, and the highest variability of zeta potential was observed at pH 5–7. A slightly decreasing trend in ZP can be observed with increasing pH in most of the samples, probably due to the increase of hydroxyl groups caused by the hydrolysis of water molecules. According to Williams and Williams [32], the increase in acidity causes the zeta potential to become less negative and even reach positive values at low pH. The resulting ZP values are noticeably low when compared with other studied kaolins [48–51]. This behavior could be related to the natural acidity of the studied kaolin aqueous suspensions, to the tendency of these kaolin samples to flocculate, and to the high viscosities measured in the samples. In fact, some authors have demonstrated a relation between zeta potential and the rheological properties of different mud-packs [52]. On the other hand, the presence of gibbsite normally implies positive surface charges in the pH range between 4 and 11 [53], and possibly affecting the magnitude of the zeta potential negativity.

Only sample VE3 showed an iso-electric point (IEP) value at  $\text{pH} \approx 6$ , while no IEPs were observed in the other samples studied (Figure 1). The IEP of kaolinite suspensions has been reported at different pH [52,54,55]. Yukselen and Kaya [56] also found showed a wide range of IEP for kaolinites, from absence to pH 6, and explained this variability as a result of kaolin composition.

Kaolinite commonly exhibits pH-dependent zeta potential due to the positively charged aluminol surface of the octahedral sheet and the edge surface sites. However, in some reported cases, it may show a largely pH-independent surface charge, suggesting cation substitution in the clay structure when exchangeable  $\text{Al}^{3+}$  releases under acidic conditions. This structural cation substitution leads to strong acid sites, where the surface charge is balanced by positive cations ( $\text{Ca}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{K}^+$ , and  $\text{Na}^+$ ) as counterions, thus explaining the slightly stable zeta potential in the 3–9 pH range [57] (Figure 1).



**Figure 1.** Zeta potentials of the studied kaolin samples with the corresponding pH, at 25 °C.

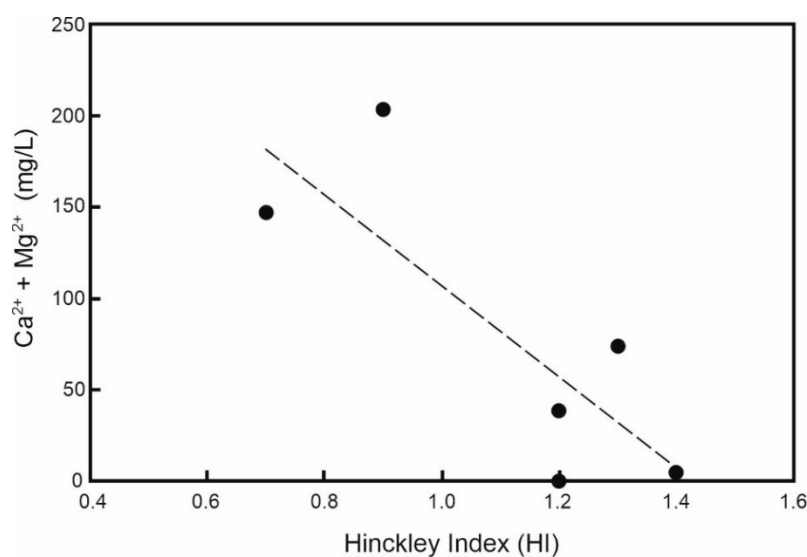
### 3.5. Desorbed Cations

Table 1 shows the contents of the kaolinite-adsorbed  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  ions released in the supernatant aqueous phases. According to Yukselen and Kaya [56], the ZP of kaolinite becomes less negative with increasing  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  concentrations in the dispersion system. Samples VE12 and VE2, which exhibited the highest values of desorbed cations, also showed the highest ZP of most of the pH values analyzed.

Normally, kaolinite does not show interlayer cation exchange. However, during primary crystallization under natural acidic conditions, strong acid sites on the edge, and broken-bond surfaces arose due to structural cation substitution, accompanied by alteration of the net proton surface charge density, leading to an increase in the affinity of cation adsorption ( $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$ ) onto the kaolinite surfaces [57,58].

We were able to correlate the calculated crystallinity of kaolinite in the studied samples with the amounts of desorbed  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  in those samples with measurable cations (Figure 2). This is probably because disordered kaolinites and those with lower HI values have more layer stacking defects and broken edges across the basal surfaces, which can provide additional sites for the adsorption of the divalent  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  cations [30,52,55].





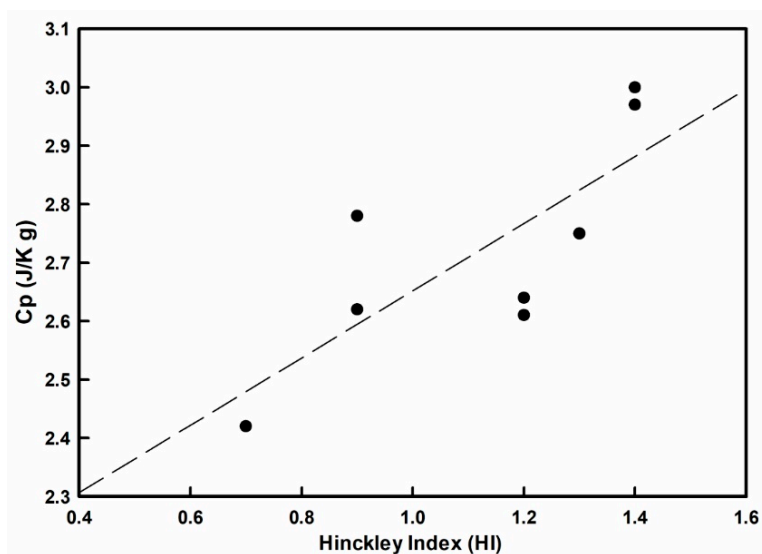
**Figure 2.** Calcium + magnesium concentrations versus Hinckley Index of the studied kaolinite samples.

### 3.6. Thermal Characteristics

Table 3 summarized the kinetic cooling parameters of the studied paste samples, including both theoretical and experimental specific heat values ( $C_p$ ), the heat retention time required to reach 32 °C ( $t_{32}$ ) and the mud temperature after 20 min ( $T_{20}$  °C), corresponding to the typical minimum application time of mud-packs, calculated by linear regression of Equation (1) ( $R^2 = 0.9999$  in all cases). Specific heat ranged from 2.42 to 3.00 J/K·g. In general, all samples showed specific heat values similar to those of some common clays evaluated for pelotherapy uses [40,59]. According to Awad et al. [29], the differences between theoretically calculated  $C_p$  values and the experimental values are normally attributed to small differences in the compositional variability at a fixed water content (50%) used in all the formulated pastes. In Egyptian sedimentary kaolins, they also found that the specific heat and the heat retention time of kaolin pastes were mainly controlled by the powder granulometric properties. The specific heat capacity ( $C_p$ ) and heat retention time ( $t_{32}$ ) of those kaolin pastes were correlated negatively with kaolinite crystallite size, which in turn was a function of both the degree of kaolinite order-disorder and its chemical variability [30]. In the Venezuelan kaolin pastes examined here, on the contrary, the degree of kaolinite order-disorder (calculated as HI) was found to correlate positively with the experimental specific heat values of the kaolin paste samples (Figure 3), and, therefore, the heterogeneous microtextural characteristics and the non-uniformity of the particle grading found in these kaolins produced a closed packing arrangement, which must be the predominant effective factor in their thermal behavior.

**Table 3.** Kinetic cooling parameters of the studied kaolin paste samples (Theor.: theoretical; Exp.: experimental).

Sample	$C_p$ (J/K·g)		$t_{32}$ °C (min)		$T_{20}$ min (°C)	
	Theor.	Exp.	Theor.	Exp.	Theor.	Exp.
VE1	2.49	2.64	25.4	24.3	34.2	34.1
VE2	2.55	2.42	24.1	23.6	33.7	33.7
VE3	2.55	2.75	24.1	23.7	33.7	33.7
VE4	2.56	2.61	23.5	23.0	33.4	33.4
VE5	2.55	2.97	27.4	25.9	34.8	34.8
VE6	2.54	3.00	26.5	26.1	34.7	34.7
VE9	2.55	2.62	27.0	26.2	34.8	34.8
VE12	2.47	2.78	27.0	26.1	35.0	35.0



**Figure 3.** Experimental specific heat capacity versus Hinckley Index.

Heat retention time (cooling from 50 to 32 °C) was found to be 23.5 to 27.4 min, with an average time of 24.8 min. The temperature corresponding to the minimal final time (20 min) of the thermal dosage administration ( $T_{20}$ ) was not below 33.7 °C. Accordingly, these values were able to assure heat transfer between the samples and the skin in normal application products. This good behavior confirmed that Venezuelan kaolins exhibit adequate thermal dosage performance and qualify them for semisolid mud packs formulations by 50% water contents for peliotherapeutic uses.

#### 4. Conclusions

Eight samples were selected from Venezuelan kaolin deposits to evaluate their sustainability for application as therapeutic muds. In general, the samples showed suitable colloidal behavior and consistent thermal dosage performance. Three samples (VE3, VE6, VE9) were particularly interesting because of their high purity (>92% kaolinite), viscosities (>1 Pa·s), adequate thermal properties (specific heat in the 2.6–3 J/K·g range, and >32 °C temperature during typical minimum topical application time), and pH range (4.4–5.9) in the optimal skin pH values for topical application. Samples showed small variations in zeta potential related to the tendency to flocculation and the high viscosities measured in the samples. There was a positive correlation between HI and the experimental specific heat values and negative correlation between HI and  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  contents. Therefore, kaolinite crystallinity influences the thermal behavior and desorption of ions on the kaolinite surface.

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